



Metal and non-metal theory test

1. Name the chemical that can absorb water.
2. Differentiate between calcination and roasting
3. What is alloy? Write its two advantages. Give two examples with composition.
4. What is anodising?
5. Draw the electron dot structure of NaCl and Al₂O₃
6. What is cinnabar and amalgam?
7. Explain the process of extraction of sodium.
8. What is thermite reaction?
9. What is amphoteric oxide?
10. Why HNO₃ is not producing H₂ gas while reacting with metals? Give exception .
11. Write down the exceptions
 - a) Metal can melt easily
 - b) Lustrous nonmetal
 - c) Non metal good conductor of electricity
 - d) Hardest non metal
 - e) Liquid non metal
12. What is enrichment of ore?
13. Name two elements can get in native state
14. What is electrovalent compound?
15. What is anode mud?
16. What is happened when Cu and Ca react with water?
17. What is the nature of oxide formed by metal and non-metal. Show it
18. Name two active metals and non-metals and how are they stored?
19. Name the reducing agent used to extract metals. Name two metals can not be extracted
20. Name the metal can not react with dilute sulphuric acid
21. Why do ionic compounds conduct electricity in molten state only?

22. Aluminium is more reactive than iron but still it is used widely?
23. Write down equations of extraction of iron
24. Explain Purification of the metal used in electric wire.
25. A metal carbonate X on reacting with an acid gives a gas which when passed through a solution Y gives the carbonate back. On the other hand, a gas G that is obtained at the anode during electrolysis of brine is passed on dry Y, it gives a compound Z, used for disinfecting drinking water. identify X, Y, G,Z

