



Metals and nonmetals

1. Assertion: the electrical conductivity and melting point of an alloy is less than that of pure metals
Reason: galvanization is a method of protecting steel and iron from rusting
2. Assertion: metals at the top of reactivity series are not formed in nature as free elements
Reason: the sulphides ore are converted into oxide by heating in the absence of air in the process of rusting
3. Assertion: metals high up in the reactivity series cannot be obtained from their compounds by heating with carbon
Reason: carbon cannot reduce the oxides of sodium, magnesium, calcium etc
4. Assertion: electrovalent compounds are generally insoluble in water
Reason: copper is higher than aluminium in the reactivity series